

1. D

Explanation: The orbital that must be filled last is the one with the highest $(n + l)$ value. For $6s$ orbital, $(n + l) = 6$; for $5d$, $(n + l) = 7$; for $4f$, $(n + l) = 7$; and for $6p$, $(n + l) = 7$. Since options (b), (c), and (d) all have the same $(n + l)$ value, we apply the rule that the one with the lower n value will be filled first. Going with these, it can be said that $6p$ orbital will be filled last.

2. C

Explanation: Based on the $(n + l)$ values calculated in No. 1, $6s$ is the 1st priority to be filled with electrons, followed by $4f$, then $5d$, and lastly, $6p$.

3. A

Explanation: The set of quantum numbers $(1, 1, 0, +\frac{1}{2})$ is impossible because an l value equal to 1 suggests that the orbital is a p orbital. Unfortunately, there is no p orbital in $n = 1$ (that's why there is no $1p^1$), hence impossible!

4. C

Explanation: Based on the mnemonic device that we are using, the correct order of orbitals to use corresponds to that in option (c). Hence, it is the correct answer.

5. A

Explanation: The only d orbital in the electron configuration of Mn is the $3d$ orbital. Therefore, $n = 3$. Furthermore, there is only one l value for d orbital, and that is 2. If we assign m_l values to the five subshells of $3d$ from -2 to $+2$, the labeled electron in the figure will fall under -1 . Therefore, the correct answer is option (a).

History Of Atoms Answer Key

Jessica J Manson



History Of Atoms Answer Key:

Whispering the Secrets of Language: An Psychological Journey through **History Of Atoms Answer Key**

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