



2. What mass in grams of P_4O_6 could be made from 8.75 g P_4 and 12.50 g O_2 ?



$$8.75g P_4 \times \frac{1 \text{ mol } P_4}{123.88g} \times \frac{1 \text{ mol } P_4O_6}{1 \text{ mol } P_4} \times \frac{219.88g}{1 \text{ mol } P_4O_6} = 15.5g P_4O_6$$

$$12.50g O_2 \times \frac{1 \text{ mol } O_2}{32.00g} \times \frac{1 \text{ mol } P_4O_6}{3 \text{ mol } O_2} \times \frac{219.88g}{1 \text{ mol } P_4O_6} = 28.63g P_4O_6$$

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LP Steffe



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