

# ATOMIC MODELS

The image displays a reading comprehension worksheet titled "Atomic Models". The central page contains a "Reading Passage" about the development of atomic models from the 19th to the 20th century. It discusses the transition from the plum pudding model to the Bohr model and the challenges posed by quantum mechanics. Below the passage are two multiple-choice questions. To the left and right of the central page are two pages of multiple-choice questions, likely representing the answer key or additional practice. In the bottom right corner, a young woman with blonde hair is smiling and holding an open book.

**Atomic Models**

**Reading Passage**

The development of atomic models has been a central pursuit in the history of science. Beginning from an understanding of the fundamental building blocks of matter, the earliest and simplest atomic models, such as Dalton's billiard ball model, depicted atoms as solid, indivisible spheres, primarily based on limited experimental data. These models focused on the basic properties of matter, such as mass and volume, but they failed to account for the internal structure of atoms.

As scientific techniques advanced in the late 19th century, scientists like J.J. Thomson proposed the plum pudding model, where electrons were embedded within a positively charged sphere. Later, Ernest Rutherford's gold foil experiments revealed that atoms contained mostly of empty space with a dense, positively charged nucleus at the center. This led to the planetary model, which depicted electrons orbiting the nucleus like planets around a sun.

However, the planetary model could not explain many observations, such as the spectral lines of elements. This challenge led to the development of quantum mechanical models in the early 20th century. These models describe electrons not as particles orbiting in fixed paths but as probability "clouds" around the nucleus. Electrons are said to be found in discrete energy levels. These models describe the probability of finding an electron in a particular region, rather than its exact position and momentum, as dictated by Heisenberg's uncertainty principle.

Experimental evidence like spectroscopy, which studies the interaction of matter with electromagnetic radiation, and scattering experiments, which probe atomic structure, have been crucial in refining atomic models. The evolution of these models illustrates the scientific process of understanding complex phenomena through continuous refinement and collaboration. The quantum mechanical model represents a significant milestone in the continuous refinement of atomic models, building on the efforts of scientists and providing a more comprehensive understanding of the structure and behavior of atoms.

Understanding these models helps us grasp how elements behave chemically, physically, and biologically, and it provides a foundation for various scientific fields, including materials science, quantum computing, and the development of new drugs. The continuous refinement of atomic models is a testament to the scientific process, driven by experimental evidence, theoretical insights, and the collaborative efforts of scientists across generations.

**Task 1: Read Atomic Models and Scientific Evidence**

Directions: Read the passage and answer the questions below.

1. Dalton's billiard ball model portrayed atoms as solid, indivisible spheres.

2.

3. The plum pudding model suggested that electrons were embedded within positively charged spheres.

4.

# Making Atomic Models Answer Key

**Susan McMurry**



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