

MC Electrochemistry

MULTIPLE CHOICE

1. Which of the following is true of an electrochemical cell?

- a) the cathode is the site of reduction
- b) the anode is negatively charged
- c) the cell voltage is independent of concentration
- d) charge is carried from one electrode to another by metal atoms passing through the solution
- e) the cathode is the site of oxidation

ANS: A

REF: AP chem 10th ed, Research & Ed Assoc

2. The standard electrode in electrochemistry involves the oxidation or reduction of which of the following?

- a) gold
- b) platinum
- c) copper
- d) magnesium
- e) hydrogen

ANS: E

REF: AP chem 10th ed, Research & Ed Assoc

3. Use the half reactions below to answer the following question:



The anode of this cell is composed of which of the following materials?

- a) Zn
- b) Cr
- c) Zn^{2+}
- d) Cr^{3+}
- e) water

ANS: A

REF: AP chem 10th ed, Research & Ed Assoc

4. One faraday of electricity (one mole of electrons) is passed through each of the following solutions.

Beaker A contains Hydrogen ions (H^{1+})

Beaker B contains Zinc ions (Zn^{2+})

Beaker C contains Silver ions (Ag^{1+})

Beaker D contains Copper(II) ions (Cu^{2+})

Which of the following answers correctly lists the amount of substance formed at one of the electrodes?

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Marco Cascella



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