

# Solution Stoichiometry

1. What volume of a 0.35M  $\text{AgNO}_3$  is required to completely react with 65 ml of a 0.29M  $\text{NaCl}$  solution?



Required: V of  $\text{AgNO}_3$

Start with the other substance

$$\frac{0.29 \text{ mol NaCl}}{1} \cdot \frac{0.055 \text{ L NaCl}}{1} \cdot \frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}}$$

then convert  
mol  $\rightarrow$  L

$$\cdot \frac{1 \text{ L AgNO}_3}{0.35 \text{ mol AgNO}_3}$$

$$= 0.0377 \text{ L} = 37.7 \text{ mL}$$

2. 37 ml of a 0.29M  $\text{KI}$  solution is required to completely react with 42 ml of a  $\text{Pb}(\text{NO}_3)_2$  solution. What was the original concentration of the lead (II) Nitrate solution?



Given:  $V_2$  37 ml  $\text{KI}$  sol  
 $M_2$  0.29M  $\text{KI}$  sol  
 $M_1$  42 ml  $\text{Pb}(\text{NO}_3)_2$

Required: original concentration of the lead

$$\frac{0.29 \text{ mol KI}}{1} \cdot \frac{0.037 \text{ L}}{1} \cdot \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{2 \text{ mol KI}} \cdot \frac{1}{0.042 \text{ L}}$$

$$= 0.1277$$

$$2(42) M_2 = 1(0.29)(37)$$

$$M_2 = 10.73 / 84$$

$$M_2 = 0.1277 \text{ M}$$

3. What mass of solid Magnesium Hydroxide can be produced if 15 mL of a 0.63M  $\text{Mg}(\text{NO}_3)_2$  solution reacts completely with excess  $\text{NaOH}$ ?



Required: mass of solid Magnesium Hydroxide if 15 ml of a 0.63M  $\text{Mg}(\text{NO}_3)_2$

Given: 15 mL  $\text{Mg}(\text{NO}_3)_2$ , 0.63M  $\text{Mg}(\text{NO}_3)_2$

$$\frac{15 \text{ mL Mg}(\text{NO}_3)_2}{1} \cdot \frac{0.63 \text{ mol Mg}(\text{NO}_3)_2}{1 \text{ L}} \cdot \frac{0.015 \text{ L}}{1} \cdot \frac{1 \text{ mol Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{NO}_3)_2}$$

$$\cdot \frac{58.32 \text{ g}}{1 \text{ mol Mg}(\text{OH})_2} = 1.65 \text{ g Mg}(\text{OH})_2$$

4. What mass of solid Calcium Phosphate can be produced if 78 ml of a 0.17M Sodium Phosphate solution is mixed with 58 ml of a 0.29M Calcium chloride solution?



Given:

78 ml  $\text{Na}_3\text{PO}_4$   
0.17M  $\text{Na}_3\text{PO}_4$

58 ml  $\text{CaCl}_2$   
0.29M  $\text{CaCl}_2$

molar mass:

$\text{Ca} - 2 \cdot 40.08$   
 $\text{P} - 2 \cdot 30.97$   
 $\text{O} - 8 \cdot 16$  } 310.18 g/mol

Required: mass (g)  $\text{Ca}_3(\text{PO}_4)_2$

$$\frac{78 \text{ mL Na}_3\text{PO}_4}{1} \cdot \frac{0.17 \text{ mol Na}_3\text{PO}_4}{1 \text{ L}} \cdot \frac{0.078 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol Na}_3\text{PO}_4}$$

$$\cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 1.925 \text{ g Ca}_3\text{PO}_4$$

From  $\text{CaCl}_2$

$$\frac{58 \text{ mL CaCl}_2}{1} \cdot \frac{0.29 \text{ mol CaCl}_2}{1 \text{ L}} \cdot \frac{0.058 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \cdot \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol}}$$

$$= 1.9389 \text{ g Ca}_3(\text{PO}_4)_2$$

# Learning Guide To Stoichiometry

**David O'Brine**



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