

# Stoichiometry

## TO GAIN AN UNDERSTANDING OF

1. Stoichiometry.
2. Limiting reagents and percent yield.

Stoichiometry is the calculation of chemical quantities from balanced equations.

## FOUR QUANTITIES IN STOICHIOMETRY CALCULATION

**particles** – the relative amounts of atoms, ions, unit formulas, or molecules in various reactants or products.

**moles** – the relative number of moles of reactants or products.

**mass** – the relative masses of reactants or products.

**volume** – the relative amounts of gaseous reactants or products.

Atoms and Mass are always conserved in chemical reactions.

## EXAMPLE OF STOICHIOMETRIC CALCULATION

Everyday Experience: the making of a peanut butter and jelly sandwich. Here is the equation:

2 slices bread + 1 serving peanut butter + 1 serving jelly  
= 1 sandwich

In this "balanced equation," we see a number of ratios: 2 slices bread/1 serving PB, 2 slices bread/1 serving jelly or 1 serving pb/1 serving jelly, etc.

## RATIOS TO DETERMINE RELATIVE AMOUNTS OF REACTANTS OR PRODUCTS

Example: If I had 12 slices of bread, how many servings of peanut butter would I need and how many sandwiches could I make? Here are the stoichiometric calculations:

$$\begin{aligned} \text{a. } 12 \cancel{\text{ slices bread}} & \left( \frac{1 \text{ serving pb}}{2 \cancel{\text{ slices bread}}} \right) = 6 \text{ servings pb.} \\ \text{b. } 12 \cancel{\text{ slices bread}} & \left( \frac{1 \text{ sandwich}}{2 \cancel{\text{ slices bread}}} \right) = 6 \text{ sandwiches} \end{aligned}$$

## BALANCED CHEMICAL EQUATION



Here are some calculations that can then be performed:

- a. If 3.50 moles of  $\text{Na}_2\text{O}$  are needed, how many moles of sodium will be required (assuming sufficient oxygen)?

$$3.50 \cancel{\text{ moles Na}_2\text{O}} \left( \frac{4 \text{ moles Na}}{2 \cancel{\text{ moles Na}_2\text{O}}} \right) = 7.00 \text{ moles Na}$$

- b. If you have 45.00 g Na and an abundant supply of oxygen, how much  $\text{Na}_2\text{O}$  should be produced?

$$45.00 \text{ g Na} \left( \frac{123.96 \text{ g Na}_2\text{O}}{91.96 \text{ g Na}} \right) = 60.66 \text{ g Na}_2\text{O}$$

- c. If you had an abundant supply of oxygen and needed to produce 12.1 moles of  $\text{Na}_2\text{O}$ , how many grams of sodium would you need?

$$12.1 \cancel{\text{ mol Na}_2\text{O}} \left( \frac{4 \cancel{\text{ mol Na}}}{2 \cancel{\text{ mol Na}_2\text{O}}} \right) \left( \frac{22.99 \text{ g}}{1 \cancel{\text{ mol Na}}} \right) = 561 \text{ g}$$

mole ratio from balanced equation      mol to g

# Lecture Notes On Stoichiometry

**Zhe Xu**



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