

### CHM1046 Exam 3 Review Questions

1. Which of the following molecules or ions is the strongest acid?
- $\text{CF}_2\text{HCO}_2\text{H}$
  - $\text{CFH}_2\text{CO}_2\text{H}$
  - $\text{CH}_3\text{CO}_2^-$
  - $\text{CH}_3\text{CO}_2\text{H}$
  - $\text{CF}_3\text{CO}_2\text{H}$

2. Assume that the following endothermic chemical reaction is at equilibrium.

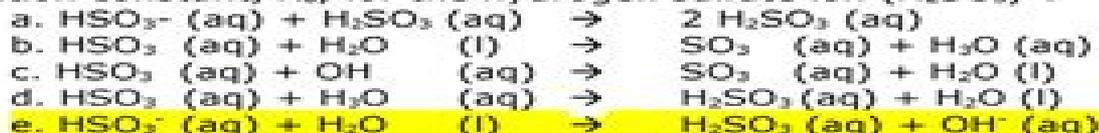


Which of the following statements is/are CORRECT?

- Increasing the concentration of  $\text{H}_2$  (g) will cause the reaction to proceed in the backward direction, increasing the equilibrium concentration of  $\text{H}_2\text{O}$ (g).
- Decreasing the temperature will cause the reaction to proceed in the forward direction, increasing the equilibrium concentration of  $\text{CO}$ (g).
- Increasing the amount of  $\text{C}$ (s) will cause the reaction to proceed in the forward direction, increasing the equilibrium concentration of  $\text{CO}$ (g).

- 1 only
- 2 only
- 3 only
- 1 and 2
- 1, 2, and 3

3. Which of the following chemical equations corresponds to the base ionization constant,  $K_b$ , for the hydrogen sulfate ion ( $\text{H}_2\text{SO}_3$ )<sup>-</sup>?



4. The equilibrium constant at a particular temperature for the reaction  $\text{H}_2 \text{ (g)} + \text{Br}_2 \text{ (g)} \rightleftharpoons 2\text{HBr (g)}$  is  $1.5 \times 10^5$ .

The value of  $K_c$  for  $\text{HBr (g)} \rightleftharpoons \frac{1}{2} \text{H}_2 \text{ (g)} + \frac{1}{2} \text{Br}_2 \text{ (g)}$  is

$$K_c = 2.58 \times 10^{-3}$$

# Miami Dade College Chm 1046 Lab Answers

**Kathleen Armour**



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### **Table of Contents Miami Dade College Chm 1046 Lab Answers**

1. Understanding the eBook Miami Dade College Chm 1046 Lab Answers
  - The Rise of Digital Reading Miami Dade College Chm 1046 Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Miami Dade College Chm 1046 Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Miami Dade College Chm 1046 Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Miami Dade College Chm 1046 Lab Answers
  - Personalized Recommendations

- Miami Dade College Chm 1046 Lab Answers User Reviews and Ratings
- Miami Dade College Chm 1046 Lab Answers and Bestseller Lists
- 5. Accessing Miami Dade College Chm 1046 Lab Answers Free and Paid eBooks
  - Miami Dade College Chm 1046 Lab Answers Public Domain eBooks
  - Miami Dade College Chm 1046 Lab Answers eBook Subscription Services
  - Miami Dade College Chm 1046 Lab Answers Budget-Friendly Options
- 6. Navigating Miami Dade College Chm 1046 Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Miami Dade College Chm 1046 Lab Answers Compatibility with Devices
  - Miami Dade College Chm 1046 Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Miami Dade College Chm 1046 Lab Answers
  - Highlighting and Note-Taking Miami Dade College Chm 1046 Lab Answers
  - Interactive Elements Miami Dade College Chm 1046 Lab Answers
- 8. Staying Engaged with Miami Dade College Chm 1046 Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Miami Dade College Chm 1046 Lab Answers
- 9. Balancing eBooks and Physical Books Miami Dade College Chm 1046 Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Miami Dade College Chm 1046 Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Miami Dade College Chm 1046 Lab Answers
  - Setting Reading Goals Miami Dade College Chm 1046 Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Miami Dade College Chm 1046 Lab Answers
  - Fact-Checking eBook Content of Miami Dade College Chm 1046 Lab Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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