

Activity A:
Small atoms

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select H (hydrogen).
- Select the ELECTRON CONFIGURATION tab.
- Click **Reset**.



Introduction: Electrons are arranged in **orbitals**, **subshells**, and **shells**. These levels of organization are shown by the boxes of the Gizmo. Each box represents an orbital. The subshells are labeled with letters (*s*, *p*, *d*, and *f*) and the shells are labeled with numbers.

Question: How are electrons arranged in elements with atomic numbers 1 through 10?

1. **Infer:** Based on its atomic number, how many electrons does a hydrogen atom have? _____
2. **Arrange:** The **Aufbau principle** states that electrons occupy the lowest-energy orbital. Click once in the **1s** box to add an electron to the only orbital in the *s* subshell of the first shell.
Click **Check**. What is the electron configuration of hydrogen? _____
3. **Arrange:** Click **Next element** to select helium. Add another electron to the **1s** orbital. The arrows represent the **spin** of the electron. What do you notice about the arrows?

The **Pauli exclusion principle** states that electrons sharing an orbital have opposite spins.

4. **Check your work:** Click **Check**. What is the electron configuration of helium? _____
5. **Arrange:** Click **Next element** and create electron configurations for lithium, beryllium, and boron. Click **Check** to check your work, and then list each configuration below:

Lithium: _____ Beryllium: _____ Boron: _____

6. **Arrange:** Click **Next element** to select carbon. Add a second electron to the first **2p** orbital.
Click **Check**. What feedback is given? _____
7. **Rearrange:** **Hund's rule** states that electrons will occupy an empty orbital when it is available in that subshell. Rearrange the electrons within the **2p** subshell and click **Check**.

Is the configuration correct now? _____

1s

Show the correct configuration in the boxes at right:

2s

2p

(Activity A continued on next page)

Element Build Explore Learning Answers

Lois Raebeck, Lawrence Wheeler



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Table of Contents Element Build Explore Learning Answers

1. Understanding the eBook Element Build Explore Learning Answers
 - The Rise of Digital Reading Element Build Explore Learning Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Element Build Explore Learning Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Element Build Explore Learning Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Element Build Explore Learning Answers
 - Personalized Recommendations
 - Element Build Explore Learning Answers User Reviews and Ratings
 - Element Build Explore Learning Answers and Bestseller Lists
5. Accessing Element Build Explore Learning Answers Free and Paid eBooks

- Element Build Explore Learning Answers Public Domain eBooks
 - Element Build Explore Learning Answers eBook Subscription Services
 - Element Build Explore Learning Answers Budget-Friendly Options
6. Navigating Element Build Explore Learning Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Element Build Explore Learning Answers Compatibility with Devices
 - Element Build Explore Learning Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Element Build Explore Learning Answers
 - Highlighting and Note-Taking Element Build Explore Learning Answers
 - Interactive Elements Element Build Explore Learning Answers
 8. Staying Engaged with Element Build Explore Learning Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Element Build Explore Learning Answers
 9. Balancing eBooks and Physical Books Element Build Explore Learning Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Element Build Explore Learning Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Element Build Explore Learning Answers
 - Setting Reading Goals Element Build Explore Learning Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Element Build Explore Learning Answers
 - Fact-Checking eBook Content of Element Build Explore Learning Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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