

Lab Conclusion - This should be written like you are turning it into an English teacher!

- 1st Paragraph: Restate the objective(s) of the lab.
(10 pts) Summarize, **in your own words**, what you did to meet the objective(s).
This should include observations that you made, data that you collected,
and calculations that you performed.
- 2nd Paragraph: Discuss anything that surprised you or any questions that you have.
(5 pts)
- 3rd Paragraph: Discuss what you learned.
(10 pts) Discuss whether or not you met the objective(s).
-If so, provide evidence (data, calculations, % error).
-If not, provide evidence(data, calculations, % error) and discuss why
not...what went wrong?

Sample Conclusion: Density Lab

The objective of the density lab was to calculate the density of water, a rubber stopper and an unknown liquid (ethanol). We also had to determine the percent error of our findings. To do so, we had to collect some data. We used the balances to determine the mass of each object. We also had to measure the volume of each using graduated cylinders. To determine the volume of the rubber stopper, we had to use water displacement. We measured the volume of water in a graduated cylinder, placed the stopper in the cylinder, and measured the new volume. The difference in the two volumes was the volume of the rubber stopper.

I was surprised to see that the density of water did calculate out to be exactly 1 g/mL. My measured mass was 46.50 grams and my measured volume was 46.5 mL, which divided out to be 1.00 g/mL. I was also surprised to see that there actually was a difference in the densities of the two liquids.

I met the objectives of the lab because I was able to determine the density of water, a rubber stopper and ethanol. I calculated the density of each by taking the mass and dividing it by the volume. Since the theoretical density of water is 1.00 g/mL, both the measured mass and measured volume were the same number, with different labels. As previously mentioned, my calculated density of water was 1.00 g/mL, with a 0% error. Since the theoretical density of ethanol is less than 1 g/mL, it made sense that the measured mass was smaller than the measured volume. My calculated density of ethanol was 0.809 g/mL, with a 2.5% error. The theoretical density of the rubber stopper was greater than 1.00 g/mL, so the measured mass was bigger than the measured volume. My calculated density for the rubber stopper was 1.29 g/mL, with a 7.5% error. My low percent errors indicate that I did well with my measurements and calculations.

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