

## LAB MANUAL WORK INSTRUCTION

**Title of Experiment:** Reaction of Sodium Sulfide and Vinegar

**Objective:** To observe the chemical reaction between sodium sulfide (sulfurized and vinegar (acetic acid) and to demonstrate the production of carbon dioxide gas.

### Pre-Lab Questions

- Write the chemical formula for sodium sulfide (sulfurized and acetic acid).
- Predict what you think will happen when these two substances are mixed. Why?
- What gas do you expect to be produced from this reaction?

### Safety Precautions

- Wear safety goggles and gloves to protect against splashes.
- Perform the experiment in a well-ventilated area to avoid inhalation of gases.
- Avoid direct skin contact with vinegar as it is an acid, though mild.

### Equipment and Materials

- 100-ml Beaker
- 50-ml Graduated Cylinder
- Baking Soda (Sodium Bicarbonate)
- Vinegar (Acetic Acid)
- Balloon
- Funnel
- Safety Goggles
- Laboratory Gloves

### Procedure

1. Place the balloon on a flat surface.
2. Using the funnel, fill the balloon with two teaspoons of baking soda.
3. Measure 50 ml of vinegar using the graduated cylinder and pour it into the beaker.
4. Carefully fit the neck of the balloon over the mouth of the beaker without getting the baking soda into the beaker.
5. Once the balloon is securely in place, lift the beaker allowing the baking soda to fall into the vinegar.
6. Observe the reaction between the baking soda and vinegar.
7. Note the speed and amount of gas produced as indicated by the inflation of the balloon.
8. Record any changes in temperature of the beaker.
9. Carefully remove the balloon from the beaker.
10. Dispose of the vinegar and baking soda solution in the sink and reuse the beaker.

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# Emerick Lab Manual

**John Augustus Smull, William P.  
Smull, Thomas Baumgardner  
Cochran, W. Harry Baker**

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