

Isotopes Practice Set

- What does "atomic number" mean?
- What does the number next to isotopes mean?
- How do you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Neon-20	Neon-22
# of protons	10	10
# of neutrons	10	12
# of electrons	10	10

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Hydrogen-1	Hydrogen-2
# of protons	1	1
# of neutrons	0	1
# of electrons	1	1

Fill in the isotopes names including atomic numbers including atomic numbers from the chart. Use your periodic table and the information provided.

	Isotope: $^{12}_6\text{C}$	Isotope: $^{16}_8\text{O}$
# of protons	6	8
# of neutrons	6	8
# of electrons	6	8

	Iron: $^{56}_{26}\text{Fe}$	Iron: $^{58}_{26}\text{Fe}$
# of protons	26	26
# of neutrons	30	32
# of electrons	26	26

	Copper: $^{63}_{29}\text{Cu}$	Copper: $^{65}_{29}\text{Cu}$
# of protons	29	29
# of neutrons	34	36
# of electrons	29	29

	Chromium: $^{52}_{24}\text{Cr}$	Chromium: $^{54}_{24}\text{Cr}$
# of protons	24	24
# of neutrons	28	30
# of electrons	24	24

Isotope Practice Set

Anthony S. Fauci



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