

# Stoichiometry

## TO GAIN AN UNDERSTANDING OF

1. Stoichiometry.
2. Limiting reagents and percent yield.

Stoichiometry is the calculation of chemical quantities from balanced equations.

## FOUR QUANTITIES IN STOICHIOMETRY CALCULATION

**particles** – the relative amounts of atoms, ions, unit formulas, or molecules in various reactants or products.

**moles** – the relative number of moles of reactants or products.

**mass** – the relative masses of reactants or products.

**volume** – the relative amounts of gaseous reactants or products.

Atoms and Mass are always conserved in chemical reactions.

## EXAMPLE OF STOICHIOMETRIC CALCULATION

Everyday Experience: the making of a peanut butter and jelly sandwich. Here is the equation:

2 slices bread + 1 serving peanut butter + 1 serving jelly  
= 1 sandwich

In this "balanced equation," we see a number of ratios: 2 slices bread/1 serving PB, 2 slices bread/1 serving jelly or 1 serving pb/1 serving jelly, etc.

## RATIOS TO DETERMINE RELATIVE AMOUNTS OF REACTANTS OR PRODUCTS

Example: If I had 12 slices of bread, how many servings of peanut butter would I need and how many sandwiches could I make? Here are the stoichiometric calculations:

$$\begin{aligned} \text{a. } 12 \cancel{\text{ slices bread}} & \left( \frac{1 \text{ serving pb}}{2 \cancel{\text{ slices bread}}} \right) = 6 \text{ servings pb.} \\ \text{b. } 12 \cancel{\text{ slices bread}} & \left( \frac{1 \text{ sandwich}}{2 \cancel{\text{ slices bread}}} \right) = 6 \text{ sandwiches} \end{aligned}$$

## BALANCED CHEMICAL EQUATION



Here are some calculations that can then be performed:

- a. If 3.50 moles of  $\text{Na}_2\text{O}$  are needed, how many moles of sodium will be required (assuming sufficient oxygen)?

$$3.50 \cancel{\text{ moles Na}_2\text{O}} \left( \frac{4 \text{ moles Na}}{2 \cancel{\text{ moles Na}_2\text{O}}} \right) = 7.00 \text{ moles Na}$$

- b. If you have 45.00 g Na and an abundant supply of oxygen, how much  $\text{Na}_2\text{O}$  should be produced?

$$45.00 \text{ g Na} \left( \frac{123.96 \text{ g Na}_2\text{O}}{91.96 \text{ g Na}} \right) = 60.66 \text{ g Na}_2\text{O}$$

- c. If you had an abundant supply of oxygen and needed to produce 12.1 moles of  $\text{Na}_2\text{O}$ , how many grams of sodium would you need?

$$12.1 \cancel{\text{ mol Na}_2\text{O}} \left( \frac{4 \cancel{\text{ mol Na}}}{2 \cancel{\text{ mol Na}_2\text{O}}} \right) \left( \frac{22.99 \text{ g}}{1 \cancel{\text{ mol Na}}} \right) = 561 \text{ g}$$

mole ratio from balanced equation      mol to g

# Lecture Notes On Stoichiometry

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## **Lecture Notes On Stoichiometry** Book Review: Unveiling the Magic of Language

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food plants of the north american indians jstor - Sep 04 2023

web tial sedentary agricultural state so far as a chronic state of war fare would permit land was set apart for farming purposes and upon it was almost invariably planted the triad of vegetables maize pumpkin or squash and beans tnis primitive agriculture was mostly pursued on and east of the mississippi in the arid in