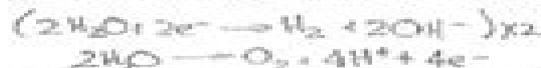


17 • Electrochemistry**ELECTROLYSIS WORKSHEET**

Standard Reduction Potential	E° (volts)
$\text{Cl}_2(\text{g}) + 2\text{e}^- \rightarrow 2\text{Cl}(\text{aq})$	+1.36
$\text{O}_2(\text{g}) + 4\text{H}^+(\text{aq}) + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}(\text{l})$	+1.23
$\text{Ag}^+(\text{aq}) + \text{e}^- \rightarrow \text{Ag}(\text{s})$	+0.80
$\text{I}_2(\text{s}) + 2\text{e}^- \rightarrow 2\text{I}^-(\text{aq})$	+0.535
$\text{Cu}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Cu}(\text{s})$	+0.337
$\text{SO}_4^{2-}(\text{aq}) + 4\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{SO}_2(\text{g}) + 2\text{H}_2\text{O}$	+0.20
$2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$ (reference electrode)	0.00
$2\text{H}_2\text{O}(\text{l}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$	-0.828
$\text{Na}^+(\text{aq}) + \text{e}^- \rightarrow \text{Na}(\text{s})$	-2.714
$\text{K}^+(\text{aq}) + \text{e}^- \rightarrow \text{K}(\text{s})$	-2.93

- All of the equations in the chart above are written as Reduction (oxidations/reductions).
- The chemicals at the upper left (Cl_2 and O_2) are the most likely to be Reduced (oxidized/reduced) and therefore the best Oxidizing agents (oxidizing agents/reducing agents).
- The chemicals at the lower right (Na and K) are the most likely to be Oxidized (oxidized/reduced) and therefore the best Reducing agents (oxidizing agents/reducing agents).
- In an electrolytic cell, the (-) electrode is negative because it has too many (too many/too few) electrons. Chemicals that come into contact with the (-) electrode will Gain (gain/lose) electrons and be Reduced (oxidized/reduced). The (-) electrode in electrolysis is called the Cathode (cathode/anode).
- Write the change that water goes through at the (-) electrode. $2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2(\text{g}) + 2\text{OH}^-(\text{aq})$
- In an electrochemical cell, the (+) electrode is positive because it has too few (too many/too few) electrons. Chemicals that come into contact with the (+) electrode will lose (gain/lose) electrons and be Oxidized (oxidized/reduced). The (+) electrode in electrolysis is called the Anode (cathode/anode).
- Write the change that water goes through at the (+) electrode. $2\text{H}_2\text{O} \rightarrow \text{O}_2 + 4\text{H}^+ + 4\text{e}^-$
- Add these two reactions together (make certain the electrons cancel) and write the overall reaction for the electrolysis of water. $2\text{H}_2\text{O} \rightarrow 2\text{H}_2(\text{g}) + \text{O}_2(\text{g})$
- We will perform this electrolysis using an aqueous solution of sodium sulfate. Both the Na^+ and H_2O will be near the (-) electrode. Which chemical is more likely to be reduced? H_2O
- Both the SO_4^{2-} and H_2O will be near the (+) electrode. Which chemical will be oxidized? H_2O



Electrolysis Experiment Lab Answer Key

SB Merriam



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Laboratory Experiments for Brown and LeMay, Chemistry, the Central Science John Henry Nelson, Kenneth C. Kemp, 1988 Laboratory Experiments for Chemistry, the Central Science, 5th Ed John Henry Nelson, 1991 **Chemistry Education in the ICT Age** Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 th th The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le M ridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings **Laboratory Manual of Inorganic Chemistry for Colleges** Lyman Churchill Newell, 1916 **Laboratory Experiments for Brown and LeMay, Chemistry, the Central Science** Nelson, John Henry Nelson, Kenneth C. Kemp, Kemp, 1981 *Energy and Water Development Appropriations for 2008* United States. Congress. House. Committee on Appropriations. Subcommittee on Energy and Water Development, 2007 A First Course in Laboratory Physics for Secondary Schools Robert Andrews Millikan, Henry Gordon Gale, Edwin Sherwood Bishop, 1914 **The Electrical Journal** ,1905 The Electrician ,1886 **Notes on Physical Laboratory Experiments in Electrical Measurements** Newell Caldwell Page, 1911 American Architect ,1904 *The American Architect and Building News* ,1903 **American Architect and the Architectural Review** ,1903 Annual Report of City Engineer ,1895 **Chemical & Metallurgical Engineering** Eugene Franz Roeber, Howard Coon Parmelee, 1920 **Report** Pennsylvania State University. Agricultural Experiment Station, 1909 *Laboratory Manual of Inorganic Chemistry* John Bernard Ekeley, 1923 Chemical News and Journal of Industrial Science ,1898 **Report of the Annual Meeting** British Association for the Advancement of Science, 1887 **Energy Conservation Update** ,1979

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