

## GAS LAW WORKSHEET 1 KEY

1. A sample of oxygen gas occupies a volume of 436.2 mL at 1.02 atm. If the temperature is held constant, what would the pressure of this gas be when the gas is compressed to 231.6 mL?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \text{since } T_1 = T_2 \quad P_1 V_1 = P_2 V_2 \quad P_2 = P_1 \frac{V_1}{V_2} = 1.02 \text{ atm} \left( \frac{436.2 \text{ mL}}{231.6 \text{ mL}} \right) = \mathbf{1.92 \text{ atm}}$$

To check with reasoning: decreased  $V$  leads to increased  $P$ ; so the volume ratio should be greater than one to yield the expected increased pressure.

2. If a gas originally occupying 6.75 L at 19.21 °C and 762.5 torr is compressed to give a pressure of 1.26 atm at 26.35 °C, what would the new volume be?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad V_2 = V_1 \left( \frac{T_2}{T_1} \right) \left( \frac{P_1}{P_2} \right) = 6.75 \text{ L} \left( \frac{299.50 \text{ K}}{292.36 \text{ K}} \right) \left( \frac{762.5 \text{ torr}}{1.26 \text{ atm}} \right) \left( \frac{1 \text{ atm}}{760 \text{ torr}} \right) = \mathbf{5.51 \text{ L}}$$

To check with reasoning: increased  $T$  leads to increased  $V$ , so the temperature ratio should be greater than one to yield the expected increase in volume. Increased  $P$  leads to decreased  $V$ , so the pressure ratio should be less than one to yield the expected decreased volume.

3. Calculate the number of grams of hydrogen sulfide gas,  $\text{H}_2\text{S}$ , in a 3.6 L container at 32.6 °C and 712 mmHg.

$$PV = \frac{g}{M} RT \quad g = \frac{PVM}{RT} = \frac{712 \text{ mmHg} (3.6 \text{ L}) 34.080 \frac{g}{\text{mol}}}{0.082058 \frac{\text{L atm}}{\text{K mol}} (305.8 \text{ K})} \left( \frac{1 \text{ atm}}{760 \text{ mmHg}} \right) = \mathbf{4.6 \text{ g H}_2\text{S}}$$

4. Calculate the density of 0.625 g of carbon dioxide at 26.32 °C and 1.03 atm.

$$PV = \frac{g}{M} RT \quad \frac{g}{V} = \frac{PM}{RT} = \frac{1.03 \text{ atm} \left( 44.010 \frac{g}{\text{mol}} \right)}{0.082058 \frac{\text{L atm}}{\text{K mol}} (299.47 \text{ K})} = \mathbf{1.84 \text{ g/L CO}_2}$$

# Gas Law Answer Key

**O García**



**Gas Law Answer Key:**

## Decoding **Gas Law Answer Key**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Gas Law Answer Key**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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