

IDEAL GAS LAW

Name: _____

Use the Ideal Gas Law below to solve the following problems.

$PV = nRT$ where P is pressure in atmospheres,
 V is volume in liters,
 n is number of moles of gas,
 R is Ideal Gas Constant $0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$,
 T is Kelvin temperature.

moles of gas: n
 volume: V
 pressure: P
 temperature: T

1. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is heated to 327 K. What is the new volume? $V = 1.12 \text{ L}$
2. A gas sample with a volume of 2.00 L at 1.00 atm and 273 K is cooled to 127 K. What is the new volume? $V = 1.50 \text{ L}$
3. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is heated to 327 K. What is the new pressure? $P = 1.12 \text{ atm}$
4. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is cooled to 127 K. What is the new pressure? $P = 0.88 \text{ atm}$
5. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is heated to 327 K. What is the new number of moles? $n = 0.041 \text{ mol}$
6. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is cooled to 127 K. What is the new number of moles? $n = 0.033 \text{ mol}$
7. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is heated to 327 K. What is the new density? $d = 1.12 \text{ g/L}$
8. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is cooled to 127 K. What is the new density? $d = 0.88 \text{ g/L}$
9. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is heated to 327 K. What is the new molar mass? $M = 29.0 \text{ g/mol}$
10. A gas sample with a volume of 1.00 L at 1.00 atm and 273 K is cooled to 127 K. What is the new molar mass? $M = 29.0 \text{ g/mol}$

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SJ Ball

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