

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Essentials

Standard Constant:

Unit	Standard Pressure
atm	1.013
mmHg	760
cmHg	76
torr	760
atm	1.013

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law:	$P_1V_1 = P_2V_2$
Charles's Law:	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law:	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law:	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law:	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Graham's Law:	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

The constant R (used in the other gas laws) depends on the unit of Pressure:

- 1 atm, $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
- 1 atm, $R = 0.0821 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$

Boyle's Law

- Temperature must always be in Kelvin ($K = ^\circ C + 273$)
- If you are given two different units of pressure, CONVERT (and)
- Create two columns before solving, initial and final, write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE (MOL) (shown)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V \propto T$

Directions: Complete each of the problems below. Show how you set up the problem and solve.

- The temperature inside my refrigerator is about 4°C. If I place it below 0°C, what would the temperature of 1.00 L of water inside my refrigerator be? (Show your work.)

Gay-Lussac's Law $P \propto T$

- A gas sample is held at a constant volume of 1.00 L and a pressure of 1.00 atm. If the temperature is increased to 2.00 times the original temperature, what is the new pressure?

Avogadro's Law $V \propto n$

- A gas sample is held at a constant pressure and temperature. If the volume is increased to 2.00 times the original volume, what is the new number of moles?

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

- A gas sample is held at a constant pressure and temperature. If the volume is increased to 2.00 times the original volume, what is the new number of moles?

Graham's Law $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

- A gas sample is held at a constant pressure and temperature. If the volume is increased to 2.00 times the original volume, what is the new number of moles?

All Gas Laws Included!



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A Guide to Teaching in the Active Learning Classroom Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment for administrators planning to create ALCs or experimenting with provisionally designed rooms and for faculty developers helping teachers transition to using these new spaces.

Cases and Materials on Oil and Gas Law, 1998 To Amend Section 541 of the Bankruptcy Code United States. Congress. Senate. Committee on the Judiciary. Subcommittee on Courts and Administrative Practice, 1989 *Fluid Mechanics Aspects of Fire and Smoke Dynamics in Enclosures* Bart Merci, Tarek Beji, 2022-10-24 This book provides essential understanding of flows in fire and smoke dynamics in enclosures covering combustion, heat transfer, and fire suppression in more detail than other introductory books. It moves from the basic equations for turbulent flows with combustion through a discussion of the structure of flames to fire and smoke plumes and their interaction with enclosure boundaries. This is then applied to fire dynamics and smoke and heat control in enclosures. This new edition provides considerably more on the fluid mechanics of the effect of water and on fire dynamics modelling using Computational Fluid Dynamics. Presents worked examples taken from practical everyday fire-related problems. Covers a broad range of topics from the basics to state-of-the-art computer simulations of fire and smoke-related fluid mechanics including the effect of water. Provides extensive treatment of the interaction of water sprays with a fire-driven flow. Contains a chapter on Computational Fluid Dynamics, the increasingly popular calculation method in the field of fire safety science. The book serves as a comprehensive guide at the undergraduate and starting researcher level.

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