

## → THE THREE STATES OF MATTER GASES, LIQUIDS & SOLIDS

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**Q.1:** Greater the temperature of the gas closer the straight line of  $p$  versus  $1/V$  to the pressure axis. Justify?

**Ans:** When the temperature of same gas increased for the number of moles, then the volume increase at the same pressure. The inverse of volumes decrease at the same pressure. So the points for the graph is above the temperature graph.

**Q.2:** Hydrogen gas diffuses faster than  $\text{CO}_2$ ?

**Ans:** As the molecular weight of  $\text{H}_2$  is lighter than  $\text{CO}_2$  and according to the Grahams law the rate of diffusion of gases is inversely proportional to the square root of their molecular weights. So hydrogen gas diffuses faster than  $\text{CO}_2$ .

**Q.3:** Why is the Boyle's law applicable only to the ideal gas?

**Ans:** Boyle's law is applicable to those gases which have no forces of attractions among the molecules. Such gases are ideal so Boyle's law of applicable to only ideal gases.

**Q.4:** Charles's law is not obeyed when the temperature is measured an Celsius Scale? Justify it?

**Ans:** The foundation stone of Charles's law is the absolute scale of temperature which starts from  $-273^\circ\text{C}$ .

**Q.5:** Dalton's law of partial pressure is only applicable to those mixtures of gases which are ideal, and they do not react with each other under the green condition? Why?

**Ans:** Dalton's law is applicable to the mixtures of gases. All the gases in the mixture should have no forces of attraction or repulsions so that they may be able to exert their own pressures independently. This is only possible when the gases are non-polar.

**Q.6:** How the value of the general gas constant "R" can be derived with the help of Avogadro's law?

**Ans:** According to Avogadro's law, the volume of one of all the ideal gases at S. T. P are  $22.4 \text{ dm}^3$ .

# Gases Section Answers

**Otis William Caldwell, Francis Day  
Curtis**



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