

Chemistry: General Equilibria and Kinetics Questions

By applying the Le Chatelier's principle, answer the following questions.

For Questions 1 to 10, the initial concentration of each species is 1.0 mol dm⁻³.

1. A mixture of nitrogen gas and hydrogen gas is contained in a closed vessel at 200°C. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
2. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
3. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
4. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
5. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
6. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
7. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
8. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
9. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.
10. A mixture of hydrogen gas and nitrogen gas is contained in a closed vessel at 200°C and 1.00 atm. The mixture is allowed to reach equilibrium at 200°C and 1.00 atm.

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