

# Gas Laws Study Guide

## Test Format:

- Multiple choice and problems

## Items that will be provided for the test (read important notes at the bottom):

- Gas law equations\*
  - The value of R.
  - Pressure conversion factors
- \* I will provide the equations on the test. You are responsible for knowing which laws the equations describe (ex: I will give you  $P_1 \times V_1 = P_2 \times V_2$  and you must know that this is Boyle's Law).

- 1) What is the name of the model we currently use to explain the behavior of gases? **kinetic molecular theory (KMT)**
- 2) Define the kinetic molecular theory (KMT). **Gases are in constant motion with elastic collisions**
- 3) The KMT makes many assumptions about the Size, motion, and energy of gas particles.
- 4) According to the KMT, do gas particles experience attractive or repulsive forces? Explain why or why not. **Neither- they are too far apart**
- 5) Completely describe the motion of gas particles according to the KMT. **Constant, Random, Straight paths, different velocity**
- 6) Define elastic collision. **No kinetic energy is lost- just transferred between particles**
- 7) What determines the kinetic energy of a particle? **Mass and velocity**
- 8) In a sample of a single gas, do all particles have the same kinetic energy? Explain why or why not. **No - they have a different velocity**
- 9) Define temperature. **Measurement of kinetic energy**
- 10) Explain why gases have a low density. **Lots of space between particles (large volume)**
- 11) Why are gases compressible? **Lots of empty space between particles**
- 12) Why do gases expand to fill a container? **No attractive or repulsive forces**
- 13) Why can gas particles flow easily past each other? **No attractive forces**
- 14) Define diffusion. **Movement of one material through another**
- 15) Define effusion. **Gas escapes through tiny openings**
- 16) Describe how diffusion depends on the mass of the particles. **Diffusion is dependent on mass, so less mass more diffusion**
- 17) Use the KMT to explain why a balloon expands as you blow air into it. **Particles are in constant motion and collide with their surroundings expanding the balloon**

## Boyle's Law: Answer the following questions about Boyle's Law.

- 18) Describe Boyle's Law in your own words. **As pressure increases, volume decreases**
- 19) What is the formula?  **$P_1V_1 = P_2V_2$**
- 20) What are the variables? What are the possible units used in the formula? **Pressure and volume, atm, kPa, mL, L**
- 21) What is constant? **temperature**



- 22) Give an example to help explain the law (you should have some type of visual of your example).
- 23) What would a graph of pressure vs. volume of a gas look like according to Boyle's Law?

## Charles' Law: Answer the following questions about Charles' Law.

- 24) Describe Charles' Law in your own words. **Increase temperature, increase volume**
- 25) What is the formula?  **$V_1/T_1 = V_2/T_2$**
- 26) What are the variables? What are the possible units used in the formula? **Temperature and volume (kelvin, mL, L)**
- 27) What is constant? **Pressure**
- 28) What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- 29) Give an example to help explain Charles's Law (you should have some type of visual of your example).



- 30) What would a graph of temperature vs. volume of a gas look like according to Charles's Law?

## Gay-Lussac's Law: Answer the following questions about Gay-Lussac's Law.

- 31) Describe Gay-Lussac's Law in your own words. **As pressure increases, temperature increases**
- 32) What is the formula?  **$P_1/T_1 = P_2/T_2$**
- 33) What are the variables? What are the possible units used in the formula? **Pressure and temperature (kelvin, atm, kPa, torr)**
- 34) What is constant? **volume**
- 35) What is very important to remember about the temperature variable when using this formula? **It must be in kelvin**
- 36) Give an example to help explain Gay-Lussac's Law (you should have some type of visual of your example).

# Gas Study Guide

**Ohio State University. Center for  
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## **Table of Contents Gas Study Guide**

1. Understanding the eBook Gas Study Guide
  - The Rise of Digital Reading Gas Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Study Guide
  - Personalized Recommendations
  - Gas Study Guide User Reviews and Ratings
  - Gas Study Guide and Bestseller Lists
5. Accessing Gas Study Guide Free and Paid eBooks
  - Gas Study Guide Public Domain eBooks
  - Gas Study Guide eBook Subscription Services
  - Gas Study Guide Budget-Friendly Options
6. Navigating Gas Study Guide eBook Formats

- ePub, PDF, MOBI, and More
  - Gas Study Guide Compatibility with Devices
  - Gas Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gas Study Guide
    - Highlighting and Note-Taking Gas Study Guide
    - Interactive Elements Gas Study Guide
  8. Staying Engaged with Gas Study Guide
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gas Study Guide
  9. Balancing eBooks and Physical Books Gas Study Guide
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gas Study Guide
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gas Study Guide
    - Setting Reading Goals Gas Study Guide
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gas Study Guide
    - Fact-Checking eBook Content of Gas Study Guide
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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