

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PV = nRT$ ", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.325 \text{ kPa} / 1 \text{ atm}$ to get $R = 8.31 \text{ kPa} \cdot L / (K \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$
$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$
$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$
$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C , what is the volume of the container that the gas is in?

$$PV = nRT$$
$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 328 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of 67°C , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$
$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 340 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$
$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

Gas Law Problems With Answers

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Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests **Questions and Answers on**

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Bank-Columbia Trust Company, 1924

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Hentz, 1982 **Problem Exercises for General Chemistry** G. Gilbert Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *C.P.A. Law Questions and Answers, 1935-1947*, American Institute of Accountants.

Board of Examiners, John Charles Teevan, 1947

Thermofluids Michael Horsley, Sherwin Keith, 1996-12-07 The two

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Pennsylvania Elementary Law, Problems and Solutions Robert Simpson Bachman, Sylvan Maxwell Flitter, 1934

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1. Understanding the eBook Gas Law Problems With Answers
 - The Rise of Digital Reading Gas Law Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Law Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Problems With Answers

- Personalized Recommendations
 - Gas Law Problems With Answers User Reviews and Ratings
 - Gas Law Problems With Answers and Bestseller Lists
5. Accessing Gas Law Problems With Answers Free and Paid eBooks
 - Gas Law Problems With Answers Public Domain eBooks
 - Gas Law Problems With Answers eBook Subscription Services
 - Gas Law Problems With Answers Budget-Friendly Options
 6. Navigating Gas Law Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Law Problems With Answers Compatibility with Devices
 - Gas Law Problems With Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Law Problems With Answers
 - Highlighting and Note-Taking Gas Law Problems With Answers
 - Interactive Elements Gas Law Problems With Answers
 8. Staying Engaged with Gas Law Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Law Problems With Answers
 9. Balancing eBooks and Physical Books Gas Law Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Law Problems With Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gas Law Problems With Answers
 - Setting Reading Goals Gas Law Problems With Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gas Law Problems With Answers

- Fact-Checking eBook Content of Gas Law Problems With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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