

13. Consider the following statements
- Atmospheric pressure is less at the top of a mountain than at sea level
  - Gases are much more compressible than solids or liquids
  - When the atmospheric pressure increases the height of the mercury column rises
- Select the correct statement
- I and II
  - II and III
  - I and III
  - I, II and III
14. Compressibility factor for  $\text{CO}_2$  at 400 K and 71.0 bar is 0.8697. The molar volume of  $\text{CO}_2$  under these conditions is
- 22.04  $\text{dm}^3$
  - 2.24  $\text{dm}^3$
  - 0.41  $\text{dm}^3$
  - 19.5  $\text{dm}^3$
15. If temperature and volume of an ideal gas is increased to twice its values, the initial pressure  $P$  becomes
- 4P
  - 2P
  - P
  - 3P
16. At identical temperature and pressure, the rate of diffusion of hydrogen gas is  $3\sqrt{3}$  times that of a hydrocarbon having molecular formula  $\text{C}_n\text{H}_{2n-2}$ . What is the value of  $n$ ?
- 8
  - 4
  - 3
  - 1
17. Equal moles of hydrogen and oxygen gases are placed in a container, with a pin-hole through which both can escape what fraction of oxygen escapes in the time required for one-half of the hydrogen to escape. (NEET phase I)
- $\frac{3}{8}$
  - $\frac{1}{2}$
  - $\frac{1}{8}$
  - $\frac{1}{4}$
18. The variation of volume  $V$ , with temperature  $T$ , keeping pressure constant is called the coefficient of thermal expansion i.e.  $\alpha = \frac{1}{V} \left( \frac{\partial V}{\partial T} \right)_P$ . For an ideal gas  $\alpha$  is equal to
- T
  - $\frac{1}{T}$
  - P
  - none of these
19. Four gases P, Q, R and S have almost same values of 'b' but their 'a' values (a, b are Vander Waals Constants) are in the order  $Q < R < S < P$ . At a particular temperature, among the four gases the most easily liquefiable one is
- P
  - Q
  - R
  - S
20. Maximum deviation from ideal gas is expected from (NEET)
- $\text{CH}_4(\text{g})$
  - $\text{NH}_3(\text{g})$
  - $\text{H}_2(\text{g})$
  - $\text{N}_2(\text{g})$
21. The units of Vander Waals constants 'b' and 'a' respectively
- $\text{mol L}^{-1}$  and  $\text{L atm}^2 \text{mol}^{-2}$
  - $\text{mol L}$  and  $\text{L atm mol}^2$
  - $\text{mol}^{-1} \text{L}$  and  $\text{L}^2 \text{atm mol}^{-2}$
  - none of these

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