



# GAS VOLUMES 1

$$\text{Volume (dm}^3\text{)} = 24 \times \text{moles}$$



The volume must be in  $\text{dm}^3$  (there are 1000  $\text{cm}^3$  in 1  $\text{dm}^3$ ).

$$\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$$

Give all answers to 3 significant figures.

- 1 Find the volume of the following gases (measured at room temperature and pressure).
- a 4.00 moles of oxygen ( $\text{O}_2$ )       $\text{volume} = 4.00 \times 24 = 96.0 \text{ dm}^3$
  - b 0.250 moles of methane ( $\text{CH}_4$ )       $\text{volume} = 0.25 \times 24 = 6.00 \text{ dm}^3$
  - c 15.0 g of argon (Ar)       $\text{moles Ar} = \frac{15}{40} = 0.375 \text{ mol}$        $\text{volume} = 0.375 \times 24 = 9.00 \text{ dm}^3$
  - d 0.220 g of carbon dioxide ( $\text{CO}_2$ )       $\text{moles CO}_2 = \frac{0.220}{44} = 0.00500 \text{ mol}$        $\text{volume} = 0.00500 \times 24 = 0.120 \text{ dm}^3$
- 2 Find the number of moles of the following gases (measured at room temperature and pressure).
- a 48.0  $\text{dm}^3$  of carbon monoxide (CO)       $\text{moles CO} = \frac{48.0}{24} = 2.00 \text{ mol}$
  - b 1.20  $\text{dm}^3$  of hydrogen ( $\text{H}_2$ )       $\text{moles H}_2 = \frac{1.20}{24} = 0.0500 \text{ mol}$
  - c 360  $\text{cm}^3$  of oxygen ( $\text{O}_2$ )       $\text{moles O}_2 = \frac{360}{24000} = 0.01500 \text{ mol}$
- 3 Find the mass of the following gases (measured at room temperature and pressure).
- a 7.20  $\text{dm}^3$  of ammonia ( $\text{NH}_3$ )       $\text{moles NH}_3 = \frac{7.20}{24} = 0.300 \text{ mol}$        $\text{mass} = 17 \times 0.300 = 5.10 \text{ g}$
  - b 480  $\text{cm}^3$  of nitrogen ( $\text{N}_2$ )       $\text{moles N}_2 = \frac{480}{24000} = 0.0200 \text{ mol}$        $\text{mass} = 28 \times 0.0200 = 0.560 \text{ g}$
  - c 100  $\text{cm}^3$  of oxygen ( $\text{O}_2$ )       $\text{moles O}_2 = \frac{100}{24000} = 0.00417 \text{ mol}$        $\text{mass} = 32 \times 0.00417 = 0.133 \text{ g}$

# Gas Variable Model 1 Answer Key

**TD Snyder**



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