

Gas Stoichiometry Practice KEY

Solve for the unknown masses and volumes using gas laws and stoichiometry

1. If you have 2 L of A and plenty of B then how many moles of C can be produced at STP?



$$\frac{2 \text{ L A}}{1} \times \frac{1 \text{ mol}}{22.4 \text{ L}} \times \frac{2 \text{ mol C}}{2 \text{ mol A}} = 0.089 \text{ mol C}$$

2. 10L of hydrogen gas with plenty of ammonia at STP will produce what volume of ammonium?



$$\frac{10 \text{ L H}_2}{1} \times \frac{1 \text{ mol}}{22.4 \text{ L}} \times \frac{2 \text{ mol NH}_4}{1 \text{ mol H}_2} \times 22.4 \text{ L} = 20 \text{ L NH}_4$$

3. 5L of hydrogen gas and 4L of oxygen gas at STP will produce what mass of water?



$$5 \text{ L} / 22.4 \text{ L} = .223 \text{ mol H}_2$$

$$4 \text{ L} / 22.4 \text{ L} = .179 \text{ mol O}_2$$

$$\frac{.223 \text{ mol H}_2}{1} \times \frac{1 \text{ mol O}_2}{2 \text{ mol H}_2} = .111 \text{ mol O}_2 \quad .111 \text{ mol O}_2 \text{ needed} < .179 \text{ mol O}_2 \text{ have. H}_2 \text{ is LR}$$

$$\frac{.223 \text{ mol H}_2}{1} \times \frac{2 \text{ mol H}_2O}{2 \text{ mol H}_2} \times \frac{18 \text{ g H}_2O}{1 \text{ mol H}_2O} = 4.01 \text{ g H}_2O$$

4. 100g of carbon and 3L of oxygen gas under the conditions of 5 atm and 50°C should produce how much carbon dioxide gas (volume)?



$$100 \text{ g} / 12 \text{ g} = 8.33 \text{ mol C}$$

$$5(3) = n(.0821)323 \Rightarrow .566 \text{ mol O}_2$$

$$\frac{8.33 \text{ mol C}}{1} \times \frac{1 \text{ mol O}_2}{1 \text{ mol C}} = 8.33 \text{ mol O}_2$$

$$8.33 \text{ mol O}_2 \text{ needed} > .566 \text{ mol O}_2 \text{ have. O}_2 \text{ is LR}$$

$$\frac{0.566 \text{ mol O}_2}{1} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol O}_2} = 0.566 \text{ mol CO}_2$$

$$5(v) = .566(.0821)323 \Rightarrow 3 \text{ L CO}_2$$

5. 20L of methane and excess oxygen gas will produce what volume of carbon dioxide and mass of water if the reactants are under 20 atm and 0°C?

$$20(20) = n(.0821)273 \Rightarrow 17.85 \text{ mol CH}_4$$

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David Callaghan



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