

Specific Heat Capacity Practice Problems

Name:

1. When 3.0 kg of water is cooled from 80.0°C to 10.0°C, how much heat energy is lost?

$$c_{p,w} = Q / m \cdot \Delta T$$

$$Q = c_{p,w} (m) \Delta T$$

$$= (4186)(3)(10 - 80)$$

$$\boxed{= -879,060 \text{ J}}$$

2. How much heat is needed to raise a 0.30 kg piece of aluminum from 30.°C to 150°C?

$$Q = c_{p,a} (m_a) (\Delta T_a)$$

$$= 899 (.3) (150 - 30)$$

$$\boxed{= +32,364 \text{ J}}$$

3. Calculate the temperature change when:

- a) 10.0 kg of water loses 232 kJ of heat.

$$232 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 232,000 \text{ J}$$

$$-232,000 \text{ J} = (4186)(10.0 \text{ kg}) \Delta T$$

$$\frac{-232,000}{4186} = \Delta T$$

$$\boxed{8.5^\circ \text{C} = \Delta T}$$

- b) 1.96 kJ of heat are added to 500. g of copper.

$$1.96 \text{ kJ} \cdot \frac{1000 \text{ J}}{1 \text{ kJ}} = 1960 \text{ J}$$

$$500 \text{ g} \cdot \frac{1 \text{ kg}}{1000 \text{ g}} = .5 \text{ kg}$$

$$1960 \text{ J} = (387)(.5 \text{ kg}) \Delta T$$

$$\frac{1960}{387} = 193.5 \Delta T$$

$$\boxed{10.1^\circ \text{C} = \Delta T}$$

4. $2.52 \times 10^4 \text{ J}$ of heat are added to 2.0 kg of mercury to reach a final temperature of 130°C. What was the initial temperature of the mercury?

$$T_f = 130^\circ \text{C}$$

$$Q = 2.52 \times 10^4$$

$$m = 2.0 \text{ kg}$$

$$c_{p,m} = 138$$

$$2.52 \times 10^4 = (138)(2)(130 - T_i)$$

$$2.52 \times 10^4 = 276(130 - T_i)$$

$$91.3 = 130 - T_i$$

$$-38.7 = -T_i$$

$$\boxed{38.7^\circ \text{C} = T_i}$$

Heat Capacity Practice Problems 4 Answers

James Hasik, Stacey Rudnick



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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