

EXP. 1 MEASUREMENT OF AN EQUILIBRIUM CONSTANT POST-LAB REPORT WORKSHEETS

Concentration of standard $\text{Na}_2\text{S}_2\text{O}_3$ solution: 0.01105

Temperature of the equilibrium mixture: 21.1 °C

Part 1 – Titration of the Aqueous Layer

You must have 3 titrations.

Volume of Aqueous Aliquots:	1	2	3
Final Burette Reading:	<u>19.61</u>	<u>36.51</u>	<u>41.68</u>
Initial Burette Reading:	<u>10.54</u>	<u>20.53</u>	<u>30.21</u>
Volume of Aliquot:	<u>9.07</u>	<u>9.98</u>	<u>11.47</u>
Titration with Sodium Thiosulfate:			
Final Burette Reading:	<u>35.98</u>	<u>37.18</u>	<u>38.93</u>
Initial Burette Reading:	<u>3.34</u>	<u>4.88</u>	<u>4.81</u>
Volume of Titer:	<u>32.64</u>	<u>32.30</u>	<u>34.12</u>
Total Concentration of Reducible Iodine [$\text{I}_2 + \text{I}_3^-$] in Aqueous Phase:	<u>0.0199</u>	<u>0.0179</u>	<u>0.01644</u>

Equilibrium Constant Lab Report Discussion

O García



Equilibrium Constant Lab Report Discussion:

Questions & Answers About Block Scheduling John Brucato, 2014-04-11 For administrators and others involved in the transition to block schedules this book provides answers to the complex and challenging questions raised by the curious and the skeptical It demonstrates how to overcome obstacles to systemic school improvements

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