

Chapter Assessment: Chemistry: Matter and Change • Chapter 4

Reviewing Vocabulary

Match each definition in Column A with the term in Column B.

Column A

Column B

- | | |
|--|---------------------------|
| 1. Radiation deflected toward the positively charged plate | a. atom |
| 2. Atoms with the same number of protons but different numbers of neutrons | b. nucleus |
| 3. High-energy radiation that has no charge and no mass | c. atomic mass |
| 4. The smallest particle of an element that retains the properties of that element | d. |
| 5. The weighted average mass of an element's isotopes | e. gamma radiation |
| 6. The center-most part of an atom where the protons and neutrons are contained | f. alpha radiation |
| 7. Radiation deflected toward the negatively charged plate | g. beta |
| 8. The rays and particles emitted by radioactive material | h. atomic |
| 9. Equal to $1/12$ the mass of a carbon-12 atom | i. |
| 10. Ray of radiation traveling from the cathode to the anode | j. Dalton's atomic theory |
| 11. Process that all matter is composed of atoms | k. cathode |
| 12. States that all matter is composed of atoms | l. |
| 13. Process whereby some substances spontaneously emit radiation | m. |
- Compare and contrast each pair of related terms.
- mass number, atomic number
 - nuclear reaction, nuclear equation

Understanding Main Ideas

Use the periodic table to identify each element described below.

- atomic number 63
- 18 protons
- 44 protons and 44 electrons
- atomic number 24
- 21 protons
- atomic number 55

In the space at the left, write true if the statement is true; if the statement is false,

change the italicized terms to make it true.

- An atom's nucleus contains its protons and electrons.
- Neutrons have no electrical charge.
- Beta particles have a charge of 2^+ .
- An alpha particle consists of two protons and two electrons.

Complete the table below.

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**Ralph M. Feather, Susan Leach
Snyder, Dale T. Hesser**



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