

Worksheet 8 - Stoichiometry calculations

1. Given the following equation: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$, show what the following molar ratios should be:

- a. $\text{C}_4\text{H}_{10} / \text{O}_2$ $2/13$
 b. O_2 / CO_2 $13/8$
 c. $\text{O}_2 / \text{H}_2\text{O}$ $13/10$
 d. $\text{C}_4\text{H}_{10} / \text{CO}_2$ $1/4$
 e. $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$ $1/5$

2. Given the following equation: $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$

How many moles of O_2 can be produced by letting 12.00 moles of KClO_3 react?

$$12.00 \text{ mol KClO}_3 \cdot \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = 18 \text{ mol O}_2$$

3. Given the following equation: $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$

(a) How many grams of KCl is produced from 2.50 g of K and excess Cl_2 . (b) How many grams of KCl is produced from 1.00 g of Cl_2 and excess K ?

$$(a) \quad 2.50 \text{ g K} \cdot \frac{1 \text{ mol K}}{39.1 \text{ g K}} \cdot \frac{2 \text{ mol KCl}}{2 \text{ mol K}} \cdot \frac{74.5 \text{ g KCl}}{1 \text{ mol KCl}} = 4.76 \text{ g KCl}$$

$$(b) \quad 1.00 \text{ g Cl}_2 \cdot \frac{1 \text{ mol Cl}_2}{70.9 \text{ g Cl}_2} \cdot \frac{2 \text{ mol KCl}}{1 \text{ mol Cl}_2} \cdot \frac{74.5 \text{ g KCl}}{1 \text{ mol KCl}} = 2.10 \text{ g KCl}$$

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Christian G. Meyer



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