

**Molar Mass and Mole Calculations Worksheet**

1. Calculate the molar mass for each of the following compounds:

- |   |  |
|---|--|
| a. NaCl = <b>58.44 g/mol</b>                      | e. $(\text{NH}_4)_3\text{PO}_4$ = <b>149.12 g/mol</b>        |
| b. $\text{Fe}_2\text{O}_3$ = <b>159.7 g/mol</b>   | f. $\text{Ca}_3(\text{PO}_4)_2$ = <b>310.18 g/mol</b>        |
| c. $\text{MgCO}_3$ = <b>84.32 g/mol</b>           | g. $\text{C}_6\text{H}_{12}\text{O}_6$ = <b>180.18 g/mol</b> |
| d. $\text{Ba}(\text{OH})_2$ = <b>171.35 g/mol</b> | h. $\text{C}_2\text{H}_5\text{N}$ = <b>79.11 g/mol</b>       |

2. Calculate the number of moles in each of the following masses:

- a. 45.0 g of acetic acid,  $\text{CH}_3\text{COOH}$  = **0.75 mol**
- b. 7.04 g of lead (II) nitrate = **0.02 mol**

3. Calculate the mass in grams of each of the following amounts:

- a. 3.00 mol of selenium oxybromide,  $\text{SeOBr}_2$  = **764.31 g**  
 $M = 78.97 + 16 + 2(79.9) = 254.77 \text{ g/mol}$      $254.77 \times 3 =$
- b. 488 mol of calcium carbonate = **48,843.92 g**  
 $M = 40.08 + 12.01 + 3(16) = 100.09 \text{ g/mol}$      $100.09 \times 488 =$

4. Calculate the number of molecules or formula units in each of the following amounts:

- a. 4.27 mol of tungsten(VI) oxide,  $\text{WO}_3$  =  **$2.57 \times 10^{24}$**   
 $4.27 \times 6.02 \times 10^{23} =$
- b. 0.989 mol of nitrobenzene  $\text{C}_6\text{H}_5\text{NO}_2$  =  **$5.95 \times 10^{23}$**   
 $0.989 \times 6.02 \times 10^{23} =$

5. Calculate the mass of each of the following quantities:

- a.  $8.39 \times 10^{23}$  molecules of fluorine = **52.96 g**  
 $8.39 \times 10^{23} / 6.02 \times 10^{23} = 1.39 \text{ mol}$      $1.39 \times 38 =$   
 $M = 2(19) = 38 \text{ g/mol}$
- b.  $6.82 \times 10^{24}$  formula units of beryllium sulfate = **1,190.33 g**  
 $M = 9.01 + 32.06 + 4(16) = 105.07$      $105.07 \times 11.33 =$   
 $6.82 \times 10^{24} / 6.02 \times 10^{23} = 11.33$

# Mole Calculations Study Guide Answer Key

**Guillaume Favre**



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### **Table of Contents Mole Calculations Study Guide Answer Key**

1. Understanding the eBook Mole Calculations Study Guide Answer Key
  - The Rise of Digital Reading Mole Calculations Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole Calculations Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Calculations Study Guide Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Calculations Study Guide Answer Key
  - Personalized Recommendations

- Mole Calculations Study Guide Answer Key User Reviews and Ratings
- Mole Calculations Study Guide Answer Key and Bestseller Lists
- 5. Accessing Mole Calculations Study Guide Answer Key Free and Paid eBooks
  - Mole Calculations Study Guide Answer Key Public Domain eBooks
  - Mole Calculations Study Guide Answer Key eBook Subscription Services
  - Mole Calculations Study Guide Answer Key Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Mole Calculations Study Guide Answer Key Compatibility with Devices
  - Mole Calculations Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Calculations Study Guide Answer Key
  - Highlighting and Note-Taking Mole Calculations Study Guide Answer Key
  - Interactive Elements Mole Calculations Study Guide Answer Key
- 8. Staying Engaged with Mole Calculations Study Guide Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Calculations Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Mole Calculations Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Calculations Study Guide Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Calculations Study Guide Answer Key
  - Setting Reading Goals Mole Calculations Study Guide Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Calculations Study Guide Answer Key
  - Fact-Checking eBook Content of Mole Calculations Study Guide Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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