

## Calorimetry and Hess's Law Post Lab Questions

**CIRCLE YOUR FINAL ANSWER and SHOW ALL WORK where applicable.**

1. Calculate the heats of combustion for the following reactions from the standard enthalpies of formation listed in Appendix 3 in your textbook:



2. From the standard enthalpies of formation, calculate  $\Delta H^\circ_{\text{rxn}}$  for the reaction



For  $\text{C}_6\text{H}_{12}(\ell)$ ,  $\Delta H_f^\circ = -151.9 \text{ kJ/mol}$ .

3. Determine the amount of heat (in kJ) given off when  $1.26 \times 10^4 \text{ g}$  of ammonia are produced according to the equation



$$\Delta H^\circ_{\text{rxn}} = -92.6 \text{ kJ/mol}$$

Assume that the reaction takes place under standard-state conditions at  $25^\circ\text{C}$ .

# Hess Law Post Lab Questions

**Tobias Bleicker**



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