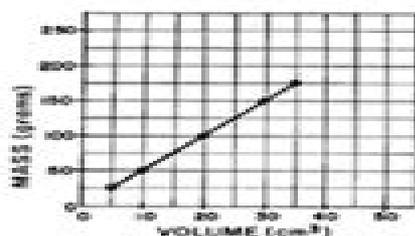


Honors Chemistry Written Final Exam Practice Test

1. Use the graph below to answer questions a-g:



a) What would be the mass of the substance if it has a volume of 15 cm³?

75 grams

b) What would be the volume of the substance if it has a mass of 175 g?

35 cm³

c) What is the density of the metal represented in the graph?

$$\text{Density} = \frac{\text{mass}}{\text{volume}} = \frac{175}{35} = 5.0 \text{ g/cm}^3$$

d) Calculate the mass of the substance above it takes up a volume 85 cm³.

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$5.0 = \frac{m}{85} \quad m = 425 \text{ g} = 430 \text{ g}$$

e) Calculate the volume of the substance above it has a mass of 325 grams.

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$5.0 = \frac{325}{V} \quad V = 65 \text{ cm}^3$$

f) Would this substance sink or float in water and why?

This substance would sink since it is more dense than water (water is 1 g/cm³)

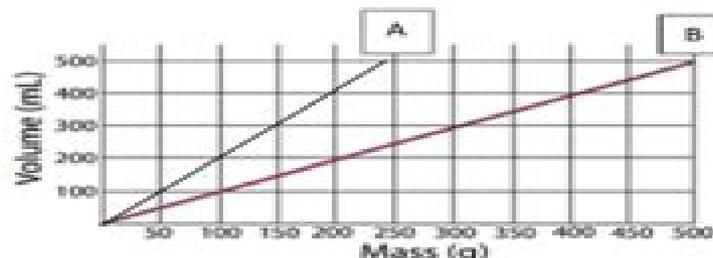
g) What does the slope of the line represent?

The density of the substance

For question h-i, use the graph to the right:

h) If the two substances were mixed together as liquids, which substance would sink to the bottom? How do you know?

Substance A is more dense since the slope of its line is steeper (you could have actually calculated density too) so substance A will sink.



i) Which would occupy a larger volume: 50 grams of substance A or 50 grams of substance B? Explain.

Since Substance B is has a lower density and you are comparing equivalent masses, Substance B would have to be taking up a larger volume.

Honors Chemistry Practice Graph Problems

DP Hallahan



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