

HOT COCOA MIX



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RJ Shavelson



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temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ C} - 25.0 \text{ C} \times q_{\text{water}}$
 $550 \text{ J} = \Delta h \times q_{\text{water}} - 550 \text{ J}$

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 for water $4.184 \text{ J/g} \times 1$ what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by 8.1 C when 207 J of heat is added 2

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 $14.73 \text{ kJ} / 14 \text{ mols} = q_{\text{rxn}} / \text{mols} \Delta t$

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 of 1.00 M NaOH at 25.0 C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0 C calculate the
 enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/ml and that

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