

Limiting Reactant Practice 2.3 - Answer Key

Objective: Use stoichiometry to identify the limiting and excess reactants.

Directions: Answer the following limiting reactant problems. Show your work!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left(\frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left(\frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.51 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left(\frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 15.6 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O₂)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 25.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left(\frac{1 \text{ mol NaOH}}{39.99 \text{ g NaOH}} \right) \left(\frac{1 \text{ mol CuSO}_4}{2 \text{ mol NaOH}} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) = 1.51 \text{ g}$$

$$25.00 \text{ g CuSO}_4 \times \left(\frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) = \boxed{15.69 \text{ g}}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO₄)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

Limiting Reactants Study Guide Answer Key

Michael Seilmaier



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28

The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane (CH₄) from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

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Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04

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