

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table



LIMITING REACTANTS

1. The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is produced in a left over?
* Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

4. Calculate moles of product * mole of reactant.

2. The reaction below produces carbon dioxide and water for a gas mixture for balloons.
 $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$

If you have 10 grams of ethene and 60 grams of oxygen, determine the limiting reactant.

3. How many grams of water produced are produced?



4. Magnesium reacts with acid to create water for drinking fountains. The reaction is:
 $Mg + 2HCl \rightarrow MgCl_2 + H_2$
If you start with 24 grams of magnesium and 2 moles of hydrochloric acid, determine the limiting reactant.

5. Calculate the grams of magnesium chloride produced.

6. How many grams of the excess reactant is produced in a left over?

5. The reaction below produces carbon dioxide and water.
 $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$
If you have 10 grams of ethene and 60 grams of oxygen, determine the limiting reactant.

6. How many grams of water produced are produced?

7. The reaction below produces carbon dioxide and water for a gas mixture for balloons.
 $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$
If you have 10 grams of ethene and 60 grams of oxygen, determine the limiting reactant.

8. Calculate the moles of product * mole of reactant.

9. Calculate the grams of product produced.

10. How many grams of the excess reactant is produced in a left over?

11. Calculate the moles of product * mole of reactant.

12. Calculate the grams of product produced.

13. How many grams of the excess reactant is produced in a left over?

ANSWER KEY INCLUDED

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EW Minium



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