

TEACHER GUIDE AND ANSWERS

Study Guide: The Structure of the Atom

Early Ideas About Matter

1. false
2. true
3. true
4. true
5. false
6. false
7. true
8. true
9. false
10. true

Defining the Atom

1. c
2. a
3. b
4. Drawing should look like a ball of chocolate chip cookie dough. The chocolate chips should be labeled with negative charge or as electrons. The dough should be labeled as evenly distributed positive charges.
5. Drawing should look like a peach with a pit. The pit should be labeled nucleus and should include labeled protons and neutrons. The outer circle of the peach should be labeled electrons.

Particle	Symbol	Location	Relative e Charge	Relative mass
6. Proton	p^+	In the nucleus	1 ⁺	1
7. Neutron	n^0	In the nucleus	0	1
8. Electron	e^-	In the space surrounding the nucleus	1 ⁻	1/1840

How Atoms Differ

1. false
2. true
3. true
4. true
5. false
6. 82 protons; 82 electrons
7. 8 protons

8. 30
9. 85
10. 104 protons; 104 electrons
11. 84 protons; 84 electrons
12. 102 protons; 102 electrons
13. 19 protons, 19 electrons, 20 neutrons
14. 14 protons, 14 electrons, 14 neutrons
15. 19 protons, 19 electrons, 21 neutrons
16. 51 protons, 51 electrons, 72 neutrons
17. The two isotopes with atomic number 19 are both isotopes of potassium.
18. $^{20}_{10}\text{Ne}$
19. ^4_2He
20. $^{133}_{55}\text{Cs}$
21. $^{238}_{92}\text{U}$
22. mass number
23. atomic number
24. d
25. a
26. d
27. 76
28. Nb
29. 190.2
30. atomic mass units
31. osmium; 76 protons, 76 electrons; niobium; 41 protons, 41 electrons.
32. Mass contribution = (mass)(percent abundance)
 $^{63}\text{X}: (62.930 \text{ amu})(69.17\%) = 43.53 \text{ amu}$
 $^{65}\text{X}: (64.928 \text{ amu})(30.83\%) = 20.02 \text{ amu}$
 Atomic mass of X = 43.53 amu + 20.02 amu = 63.55 amu
 The element is copper.
33. Mass contribution = (mass)(percent abundance)
 $^{35}\text{X}: (34.969 \text{ amu})(75.77\%) = 26.50 \text{ amu}$
 $^{37}\text{X}: (36.966 \text{ amu})(24.23\%) = 8.957 \text{ amu}$
 Atomic mass of X = 26.50 amu + 8.957 amu = 35.46 amu
 The element is chlorine.

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