

Chapter 17: Chemical Equilibrium

Worksheet: 17.1 (1 of 2)

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Section 1

1. Consider the reaction:



At equilibrium, the concentration of NO_2 is 0.040 M and the concentration of N_2O_4 is 0.090 M .

Calculate the equilibrium constant K_c for this reaction.

2. For the reaction $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$, the equilibrium constant K_c is 0.50 at 400°C .

Calculate the equilibrium constant K_p for this reaction at 400°C .

3. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

4. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

5. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

6. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

Section 2

7. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

8. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

9. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

10. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

11. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

12. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

13. The equilibrium constant K_c for the reaction $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ is 4.0×10^{-4} at 1500°C .

Calculate the equilibrium constant K_p for this reaction at 1500°C .

Elements Of Language First Course Answer Key

L Darling-Hammond



Elements Of Language First Course Answer Key:

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