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**MATHEMATICS
HIGHER LEVEL
PAPER 1**

Wednesday 12 November 2014 (afternoon)

2 hours

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 temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ j/g} \cdot \text{c} \times 110 \text{ g} \times 26.6 \text{ c} - 25.0 \text{ c} \times q_{\text{water}}$
 $550 \text{ j} = \Delta h \times q_{\text{water}} = 550 \text{ j}$

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 $14.73 \text{ kJ} = \Delta h_{\text{rxn}} = q_{\text{rxn}} = q_{\text{calorimeter}} = m c \Delta t$

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