

## FORMULAS FOR IONIC COMPOUNDS

1. Write the chemical formula for the ionic compounds made from the following elements.

Elements	metal ion	non-metal ion	Chemical Formula	Name
potassium and iodine	$K^{+1}$	$I^{-1}$	KI	Potassium Iodide
barium and chlorine	$Ba^{2+}$	$Cl^{-1}$	$BaCl_2$	Barium Chloride
sodium and oxygen	$Na^{+1}$	$O^{2-}$	$Na_2O$	sodium Oxide
aluminum and nitrogen	$Al^{3+}$	$N^{3-}$	AlN	Aluminum Nitride
lithium and sulfur	$Li^{+1}$	$S^{2-}$	$Li_2S$	Lithium Sulfide
magnesium and oxygen	$Mg^{+2}$	$O^{2-}$	MgO	Magnesium Oxide
calcium and chlorine	$Ca^{2+}$	$Cl^{-1}$	$CaCl_2$	calcium chloride
boron and chlorine	$B^{3+}$	$Cl^{-1}$	$BCl_3$	boron chloride
beryllium and sulfur	$Be^{2+}$	$S^{2-}$	BeS	beryllium sulfide

2. Complete the chart below and write the ionic formulas to complete it. A few have been done for you.

	F	Cl	$O^{2-}$	$S^{2-}$	$N^{3-}$	$P^{3-}$
$Li^{+}$	$LiF$	$LiCl$	$Li_2O$	$Li_2S$	$Li_3N$	$Li_3P$
$Na^{+}$	$NaF$	$NaCl$	$Na_2O$	$Na_2S$	$Na_3N$	$Na_3P$
$K^{+}$	$KF$	$KCl$	$K_2O$	$K_2S$	$K_3N$	$K_3P$
$Be^{2+}$	$BeF_2$	$BeCl_2$	$BeO$	$BeS$	$Be_3N_2$	$Be_3P_2$
$Mg^{2+}$	$MgF_2$	$MgCl_2$	$MgO$	$MgS$	$Mg_3N_2$	$Mg_3P_2$
$Ca^{2+}$	$CaF_2$	$CaCl_2$	$CaO$	$CaS$	$Ca_3N_2$	$Ca_3P_2$
$Al^{3+}$	$AlF_3$	$AlCl_3$	$Al_2O_3$	$Al_2S_3$	$AlN$	$AlP$
$Ga^{3+}$	$GaF_3$	$GaCl_3$	$Ga_2O_3$	$Ga_2S_3$	$GaN$	$GaP$
$B^{3+}$	$BF_3$	$BCl_3$	$B_2O_3$	$B_2S_3$	$BN$	$BP$

# Ionic Formula Writing Kit Lab Answer Key

**Yicheng Fang**



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## **Table of Contents Ionic Formula Writing Kit Lab Answer Key**

1. Understanding the eBook Ionic Formula Writing Kit Lab Answer Key
  - The Rise of Digital Reading Ionic Formula Writing Kit Lab Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Formula Writing Kit Lab Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Formula Writing Kit Lab Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Formula Writing Kit Lab Answer Key
  - Personalized Recommendations
  - Ionic Formula Writing Kit Lab Answer Key User Reviews and Ratings
  - Ionic Formula Writing Kit Lab Answer Key and Bestseller Lists
5. Accessing Ionic Formula Writing Kit Lab Answer Key Free and Paid eBooks
  - Ionic Formula Writing Kit Lab Answer Key Public Domain eBooks
  - Ionic Formula Writing Kit Lab Answer Key eBook Subscription Services

- Ionic Formula Writing Kit Lab Answer Key Budget-Friendly Options
- 6. Navigating Ionic Formula Writing Kit Lab Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic Formula Writing Kit Lab Answer Key Compatibility with Devices
  - Ionic Formula Writing Kit Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic Formula Writing Kit Lab Answer Key
  - Highlighting and Note-Taking Ionic Formula Writing Kit Lab Answer Key
  - Interactive Elements Ionic Formula Writing Kit Lab Answer Key
- 8. Staying Engaged with Ionic Formula Writing Kit Lab Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ionic Formula Writing Kit Lab Answer Key
- 9. Balancing eBooks and Physical Books Ionic Formula Writing Kit Lab Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic Formula Writing Kit Lab Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Formula Writing Kit Lab Answer Key
  - Setting Reading Goals Ionic Formula Writing Kit Lab Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Formula Writing Kit Lab Answer Key
  - Fact-Checking eBook Content of Ionic Formula Writing Kit Lab Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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