

### The Nucleus, A Chemist's View

The nucleus of an atom is not always stable. Radioactive decay can occur, where components of a nucleus leave the atom, turning the atom into a different isotope, discussed heavily earlier in the class. Instability is more common than not, over the range of all the different kinds of nuclei possible. Out of the more than 2000 known nuclides (forms of nucleus), only 279 are considered stable. It is also important to note that any nucleus with more than 83 protons is by definition unstable. Another general rule is that for lighter nuclides, the isotopes are the most stable when the neutron:proton ratio is close to one, whereas for heavier nuclides the stability is enhanced when the neutron:proton ratio is more than one. For reasons of symmetry it has also been discovered that proton levels of 2, 8, 20, 28, 50, 82, and 126 are even though the last number obviously has far more protons than should be present in a stable nuclide.

The number of neutrons and protons are seen in the notation used for radioactive decay problems. For example, a carbon atom with 6 neutrons would be noted like this:



with the number of protons represented by the bottom number while the top number is the total of both protons and neutrons (so the number of neutrons is the top number minus the bottom number). Oftentimes the bottom number is dropped, as the number of protons follows naturally from the elemental symbol (each element can only have one particular number of protons, or else it transforms into a different element). There is a zone of stability when it comes to ratios, as can be seen on the following chart.

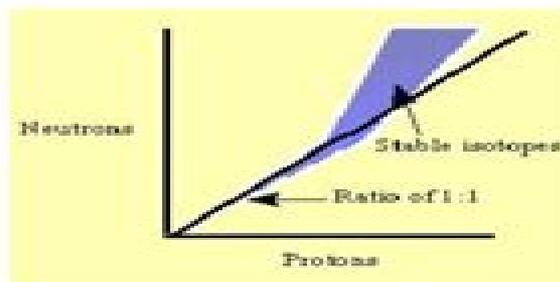


Figure 1. Zone of Stability

There are several different types of radioactive decay. The first is Alpha emissions, the alpha particle is equivalent to a helium nucleus, and this occurs only with extremely heavy nuclei. For example,



Beta emissions release beta particles instead, which are electrons emitted by a nucleus after a neutron has been converted into a proton and electron. This process is represented as follows,

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**Camilla Rothe**



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