

Use the table of standard reduction potentials from your textbook to complete the following table. The zinc-copper Daniell cell is given as an example.

Electrodes	Conc. of Each Cation	Half-Reactions	Cathode or Anode?	E° (table)	E°_{cell} (calc'd)	E°_{cell} (measured)
A. Determine the E° for the Standard Zn-Cu Galvanic Cell						
Zn	1.0 M Zn^{2+}	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	+1.10 V	0.93 V
Cu	1.0 M Cu^{2+}	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Cathode	+0.34 V		
B. Determine the E° for the Cu-Zn Cell at Non-standard Concentrations						
Zn	1.0 M Zn^{2+}	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	_____ V	0.95 V
Cu	0.5 M Cu^{2+}	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Cathode	+0.34 V		
Zn	1.0 M Zn^{2+}	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	_____ V	0.96 V
Cu	0.1 M Cu^{2+}	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Cathode	+0.34 V		
C. Determine the E° for Zn-Mg and Cu-Mg Galvanic Cells						
Zn	1.0 M Zn^{2+}	$\text{Zn}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Zn}_{\text{s}}$	Anode	-0.76 V	_____ V	0.82 V
Mg	1.0 M Mg^{2+}	$\text{Mg}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Mg}_{\text{s}}$	Cathode	-2.38 V		
Cu	1.0 M Cu^{2+}	$\text{Cu}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Cu}_{\text{s}}$	Anode	+0.34 V	_____ V	1.10 V
Mg	1.0 M Mg^{2+}	$\text{Mg}^{2+}_{\text{aq}} + 2e^- \rightarrow \text{Mg}_{\text{s}}$	Cathode	-2.38 V		

*You must use the Nernst equation to calculate this E_{cell} voltage. (See Equation 8.)

Galvanic Cell Experiment Report

A Loxley



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Table of Contents Galvanic Cell Experiment Report

1. Understanding the eBook Galvanic Cell Experiment Report
 - The Rise of Digital Reading Galvanic Cell Experiment Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Galvanic Cell Experiment Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Galvanic Cell Experiment Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Galvanic Cell Experiment Report
 - Personalized Recommendations
 - Galvanic Cell Experiment Report User Reviews and Ratings
 - Galvanic Cell Experiment Report and Bestseller Lists
5. Accessing Galvanic Cell Experiment Report Free and Paid eBooks
 - Galvanic Cell Experiment Report Public Domain eBooks
 - Galvanic Cell Experiment Report eBook Subscription Services
 - Galvanic Cell Experiment Report Budget-Friendly Options
6. Navigating Galvanic Cell Experiment Report eBook Formats

- ePub, PDF, MOBI, and More
 - Galvanic Cell Experiment Report Compatibility with Devices
 - Galvanic Cell Experiment Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Galvanic Cell Experiment Report
 - Highlighting and Note-Taking Galvanic Cell Experiment Report
 - Interactive Elements Galvanic Cell Experiment Report
 8. Staying Engaged with Galvanic Cell Experiment Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Galvanic Cell Experiment Report
 9. Balancing eBooks and Physical Books Galvanic Cell Experiment Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Galvanic Cell Experiment Report
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Galvanic Cell Experiment Report
 - Setting Reading Goals Galvanic Cell Experiment Report
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Galvanic Cell Experiment Report
 - Fact-Checking eBook Content of Galvanic Cell Experiment Report
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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