

Name: \_\_\_\_\_ **KEY** \_\_\_\_\_ Period: \_\_\_\_\_ Date: \_\_\_\_\_

## Determining Molecular Polarity

Molecules are either POLAR or NON-POLAR. This is determined by:

- If all of the bonds are ionic → the compound is just "ionic"
- If all of the bonds are non-polar covalent → then the molecule is non-polar
- If one of the bonds is polar → then the molecule is polar
- If there is more than one polar bond, then look at symmetry
  - If the shape linear, tetrahedron, or trigonal planar and all the bonds are the same, then it had symmetry and is non-polar.
  - If the shape is a pyramid, linear, or bent, then the molecules in NOT symmetrical and is a polar molecule

NPC = Non-Polar Covalent  
 PC = Polar Covalent  
 I = Ionic

Fill in the following table:

Compound	Lewis Dot Diagram	Bond Type (NPC/PC/I)	Shape	Symmetry (Y/N)	Molecular Polarity (P/NP)
H <sub>2</sub>	H — H	H—H (NPC)	Linear	YES	NP
N <sub>2</sub>	: N ≡ N :	N—N (NPC)	Linear	YES	NP
O <sub>2</sub>	: $\overset{\cdot\cdot}{\text{O}} = \overset{\cdot\cdot}{\text{O}}$ :	O—O (NPC)	Linear	YES	NP
CaCl <sub>2</sub>	: $\overset{\cdot\cdot}{\text{Cl}}$ : Ca : $\overset{\cdot\cdot}{\text{Cl}}$ :	(I)	No Shape (Ionic)	NONE	IONIC
CO	: $\overset{\cdot\cdot}{\text{C}} \equiv \overset{\cdot\cdot}{\text{O}}$ :	C—O (PC)	Linear	NO	P

# Molecular Shape And Polarity Study Guide Answers

**R Barnett**



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